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Practice Problems 1.
Molar solutions. a.
Describe how you
would prepare 1 L of a
1 M solution of sodium
chloride. The gram
formula weight of

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sodium chloride is 58.44 g/mol. Answer: To make a 1 M solution of sodium chloride, dissolve 58.44 g sodium chloride in 500 mL water in a 1000-mL volumetric flask. When all the solid is dissolved and the solution is at room temperature, dilute to the mark and invert the flask several times to mix.

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The measurement of how large a mixture can get before it becomes a solution. The measurement of how much solute can be dissolved in a liter of solvent. The proportion of solute to solvent in a mixture. None of the Above.

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Name _____ MULTIPLE

CHOICE. Choose the

one alternative that

best completes the

statement or answers

the question. 1)

Formation of solutions

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where the process is endothermic can be spontaneous provided that _____. A) the solvent is a gas and the solute is a solid

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Practice Problems:
Solutions (Answer Key)
What mass of solute is
needed to prepare
each of the following

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solutions? a. 1.00 L of
0.125 M K_2SO_4 21.8
g K_2SO_4 b. 375 mL of
0.015 M NaF 0.24 g
NaF c. 500 mL of 0.350
M $C_6H_{12}O_6$ 31.5 g
 $C_6H_{12}O_6$; Calculate
the molarity of each of
the following solutions:

Practice Problems: Solutions - Department of Chemistry

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and adding more solute, then cooling the solution.

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